

**(12) UK Patent Application (19) GB (11) 2 330 137 (13) A**

(43) Date of A Publication **14.04.1999**

(21) Application No **9820167.6**

(22) Date of Filing **17.09.1998**

(30) Priority Data  
(31) **08946744** (32) **09.10.1997** (33) **US**

(71) Applicant(s)  
**MBT Holding AG**  
(Incorporated in Switzerland)  
Vulkanstrasse 110, CH-8048 Zürich, Switzerland

(72) Inventor(s)  
**Stephen A Farrington**

(74) Agent and/or Address for Service  
**P B Brown**  
Feb Mbt, Albany House, Swinton Hall Road, Swinton,  
**MANCHESTER, M27 4DT, United Kingdom**

(51) INT CL<sup>6</sup>  
**C04B 9/04 28/34 // C04B 111:10 111:72**

(52) UK CL (Edition Q )  
**C1H HBX H660 H718 H721 H727 H740 H742 H795**

(56) Documents Cited  
**WO 97/34848 A1 WO 97/21639 A1 US 5595597 A**

(58) Field of Search  
UK CL (Edition P ) **C1H HBX**  
INT CL<sup>6</sup> **C04B 9/00 9/04 12/00 12/02 28/00 28/10**  
**28/34**  
**ONLINE: EDOC, JAPIO, WPI**

(54) Abstract Title  
**Cementitious materials**

(57) A method for making a cementitious material comprises forming a mixture of water and active ingredients consisting essentially of i) a compound capable of providing, in aqueous solution, magnesium ions, ii) a potassium phosphate compound and iii) optionally an additional phosphate source; wherein forming said mixture is substantially in the absence of ammonia or an ammonium ion-containing compound. A method for repairing a flaw in a cementitious structure comprises applying to the flaw the cementitious binder thus prepared.

**GB 2 330 137 A**

## RAPID SETTING, HIGH EARLY STRENGTH BINDERS

### TECHNICAL FIELD

5

The present invention is directed to cementitious binders. More particularly, the present invention is directed to rapid-setting, high early strength binders for use as cementitious repair materials, comprising phosphate cements.

10

### BACKGROUND OF THE INVENTION

15

High early strength cementitious binders have been used in the past to repair cementitious structures, or as components of cementitious compositions to form structures. In one class of these materials, ammonium phosphates are used to prepare binder materials in reaction with magnesium ion-providing materials, such as MgO. International Patent Publication WO 96/35647 discloses a mechanism by which potassium struvite was prepared from an ammonium phosphate reactant. Some have proposed the use of other metal phosphates in conjunction with or as a replacement of ammonium phosphate, such as phosphate acid salts of Al, Ca, Mg or Na.

20

USPN 2,522,548 discloses the production of refractory molds for metal casting made with phosphate binders produced from a metal oxide or hydroxide (Ca, Mg or Zn), a primary metal phosphate (Mg, Ca, Zn, Mn, Al or Fe), and a primary alkali metal phosphate (Na or K). The binder may be combined with fillers and retarders.

25

USPN 4,444,594 discloses acid-cured inorganic compositions suitable for a binder material compatible with mineral wool, for application in ceiling tiles. The binder is produced by reacting MgO, an acid phosphate, chloride or sulfate salt, an amino alcohol and water. Fillers can be added to the binder. The acid phosphate can be ammonium phosphate (preferred), sodium phosphate and potassium phosphate.

30

One problem encountered with the use of conventional repair binders or mortars is that they usually contain ammonium phosphate hydrate as a component, for reaction with an alkaline earth oxide or hydroxide such as MgO or MgOH in water to form struvite. The reaction of the ammonium phosphate salt with the alkaline earth metal drives off ammonia as a gas, causing foaming and creating an unpleasant environmental odor. The use of such binders or cements indoors is severely limited.

A further problem with conventional materials is that the reaction between the alkaline earth metal ion and the ammonium phosphate salt is so fast that the binder sets almost immediately, and limits the ability of the workman to "finish" the product in place. Setting retarders are therefore needed to permit working of the mixture to effectively treat the structure in need of repair.

### SUMMARY OF THE INVENTION

15

It is therefore an object of the invention to provide a high early strength binder based on phosphate cement, which avoids the evolution of ammonia gas in its preparation.

20

It is a further object of the present invention to provide a high early strength binder that has a set time conducive to workability of the binder and finishing operations.

25

The present invention is directed to a process for preparing a cementitious binder useful in a high strength, quick-setting mortar, whose utility includes a repair material for cementitious structures. The binding system is based on the formation of potassium struvite, by reacting a source of magnesium with potassium phosphate and water. Although still quick setting, the set time of this binder is slower than the set time of struvite. Inert fillers are included in mortars containing this material, and retarders may be included in the formulation to lengthen the workability time before set.

30

A significant advantage of the invention is the elimination of the conventionally-used ammonia or ammonium ion-containing materials which are used in the industry for quick-setting mortars based on struvite. Avoidance of the evolution of ammonia in the use of the mortars greatly expands their applicability.

5

The present invention provides a method for making a cementitious material comprising:

forming a mixture of water and active ingredients consisting essentially of i) a compound capable of providing, in aqueous solution, magnesium ions, ii) a potassium phosphate compound, and iii) optionally an additional phosphate source;

10

wherein said mixture is formed substantially in the absence of ammonia or an ammonium ion-containing compound. In one embodiment, the compound capable of providing magnesium ions is  $\text{MgO}$ , and the potassium phosphate compound is  $\text{KH}_2\text{PO}_4$ .

15

The present invention further provides a method for repairing a flaw in a cementitious structure, comprising applying to the flaw a cementitious binder prepared by

forming a mixture of water and active ingredients consisting essentially of i) a compound capable of providing, in aqueous solution, magnesium ions, ii) a potassium phosphate compound, and iii) optionally an additional phosphate source;

20

wherein said mixture is formed substantially in the absence of ammonia or an ammonium ion-containing compound.

25

In one embodiment, the cementitious binder of the present invention is formed in a mixture to prepare a mortar, wherein forming the mixture includes adding a substantially inert filler.

30

The present invention further provides a method for repairing a flaw in a cementitious structure, comprising applying to the flaw a mortar prepared by

forming a mixture of an inert filler, water, and active ingredients consisting

essentially of i) a compound capable of providing, in aqueous solution, magnesium ions, ii) a potassium phosphate compound, and iii) optionally an additional phosphate source;

5 wherein forming said mixture is substantially in the absence of ammonia or an ammonium ion-containing compound.

### DETAILED DESCRIPTION OF THE INVENTION

10 The present invention provides a method and binder formulations for preparing high strength, quick setting cementitious compositions, useful for concretes, flooring materials, grouts, and repair or patch materials for cementitious structures and the like.

The method and compositions of the present invention are useful in construction materials which are mixed with water in close proximity to the place where they are to be cast or placed.

15

It is desirable that the materials of the present invention be usable indoors, which limits the possible components of the formulation used to prepare them. For example, the ammonium ion-containing formulations of the prior art which have been used as repair materials evolve ammonia gas during the water-initiated reaction, and thus generally cannot be used indoors.

20

The present invention eliminates the generation of ammonia gas by providing a formulation and method which permits the formation of a high-strength cementitious binder, substantially in the absence of ammonia or an ammonium ion-containing compound.

25

The binding system of present invention is based on the formation of potassium struvite ( $\text{MgKPO}_4 \cdot 6\text{H}_2\text{O}$ ), by reacting a source of magnesium with a potassium phosphate compound such as  $\text{KH}_2\text{PO}_4$  and water. The method includes forming a mixture of water and active ingredients consisting essentially of i) a compound capable of providing, in aqueous solution, magnesium ions, and ii) a potassium phosphate

30

compound; wherein the mixture is formed substantially in the absence of ammonia or an ammonium ion-containing compound. In a preferred embodiment, the compound capable of providing magnesium ions is  $\text{MgO}$ , and the potassium phosphate compound is  $\text{KH}_2\text{PO}_4$ .

5

Other sources of magnesium ions that can be used in the present method include but are not limited to magnesium hydroxide, magnesium phosphate hydrate, magnesium acetate, magnesium sulfate, magnesium thiosulfate, magnesium nitrite, magnesium thiocyanate, magnesium phosphate (monobasic), magnesium bromide,  
10 magnesium citrate, magnesium nitrate, and compatible mixtures thereof.

15

Based on active ingredients (*i.e.* magnesium source, potassium phosphate compound, retarder, additional phosphate source, and water), the formulation preferably contains about 10 to about 35 percent by weight of the magnesium source compound(s), preferably about 13 to about 30 percent by weight, and about 30 to about 60 percent by weight of the potassium phosphate compound, preferably about 30 to about 56 percent by weight. Water is generally added in an amount of about 15 to about 40 percent by weight, preferably about 16 to about 39 percent by weight, in order to initiate the reaction between the other components of the formulation.

20

The cementitious binder, thus prepared, can be used as a grout, or as a patch or repair material, by applying the binder, before setting, to the target surface.

25

Although still quick-setting, the set time of this binder is slower than the set time of struvite. This permits the binder to be worked into a cast, mold, patch or other configuration while in a plastic state.

30

If desired, the set time can be further slowed by the optional addition of a retarder to the mixture, generally in the amount of up to 3 percent by weight. Examples of suitable retarders include, but are not limited to, oxy-boron compounds, polyphosphonic acid, salts of polyphosphonic acid, carboxylic acids, salts of

carboxylic acids, polycarboxylic acids, salts of polycarboxylic acids, hydroxycarboxylic acids, salts of hydroxycarboxylic acids, alkali metal salts of halogens, nitrates, nitrites, or sulfates, alkali metal hydroxides, alkali metal carbonates, and mixtures thereof.

5

Further examples of suitable retarders for the inventive cementitious binder include, but are not limited to boric acid, alkali metal borates such as sodium borate, sodium tetraborate decahydrate, trialkoxyborates such as trimethylborate, triarylborates, nitrilotris(methylene) tris(phosphonic)acid and its pentasodium salt, 10 citric acid, sodium citrate (monobasic or dibasic), trimellitic acid (hydrous), nitrilotriacetic acid trisodium salt (monohydrate), sodium chloride, potassium chloride, sodium silico-fluoride, sodium nitrate, potassium nitrate, sodium nitrite, potassium nitrite, sodium sulfate, potassium sulfate, sodium chloride, potassium chloride, sodium carbonate, potassium carbonate, and the like.

15

Other phosphate sources for the binder can optionally be utilized in addition to, or to replace a portion of, the potassium phosphate. The additional phosphate sources may further function as a dispersant for the binder and/or an accelerator for the reaction, and can contribute to strength development of the material. Examples of 20 such phosphate sources include, but are not limited to, sodium tripolyphosphate, potassium tripolyphosphate, and sodium potassium tripolyphosphate. The additional phosphate source can be added to the other active ingredients in the mixture in an amount of up to about 20 percent by weight, preferably up to about 15% by weight.

25

Examples of the present inventive cementitious binder and method were conducted in comparison to prior art ammonium ion-containing formulations. A magnesium source, phosphate source, water and optionally a retarder were mixed, placed in a mold and allowed to set. Specific binder mixture components and their amounts, are set forth in Table 1 below.

30

**Table 1**

<u>Example</u>	<u>Binder Mixture Formulation</u>
5	C 1.
	MgO(38%) - $\text{NH}_4\text{H}_2\text{PO}_4$ (43%) - $\text{H}_2\text{O}$ (19%)
	2.
	MgO(28%) - $\text{KH}_2\text{PO}_4$ (55%) - $\text{H}_2\text{O}$ (17%)
10	C 3.
	MgO(33%) - $\text{NH}_4\text{H}_2\text{PO}_4$ (32%) - $\text{K}_2\text{CO}_3$ (18%) - $\text{H}_2\text{O}$ (17%)
	C 4.
	MgO(39%) - $\text{NH}_4\text{H}_2\text{PO}_4$ (44%) - $\text{H}_2\text{O}$ (17%)
	5.
	MgO(28%) - $\text{KH}_2\text{PO}_4$ (55%) - $\text{H}_2\text{O}$ (17%)
15	C 6.
	MgO(33%) - $\text{NH}_4\text{H}_2\text{PO}_4$ (32%) - $\text{K}_2\text{CO}_3$ (18%) - $\text{H}_2\text{O}$ (17%)
	C 7.
	MgO(39%) - $\text{NH}_4\text{H}_2\text{PO}_4$ (44%) - $\text{H}_2\text{O}$ (17%)
	8.
	MgO(28%) - $\text{KH}_2\text{PO}_4$ (55%) - $\text{H}_2\text{O}$ (17%)
20	C 9.
	MgO(33%) - $\text{NH}_4\text{H}_2\text{PO}_4$ (32%) - $\text{K}_2\text{CO}_3$ (18%) - $\text{H}_2\text{O}$ (17%)

Mixes 4-6 contained 0.2 wt. % boric acid.

Mixes 7-9 contained 0.5 wt. % boric acid.

25

The fluidities of the initial mixtures were evaluated, the setting times measured using the Vicat apparatus described in ASTM C-191, and the compressive strengths of the molded articles tested after 24 hours as described in ASTM C-109. Results are set forth in Table 2, below.



**Table 2**

	<u>Example</u>	<u>Fluidity</u>	<u>Setting time (minutes)</u>	<u>24 hour compressive strength (MPa)</u>
5	C 1.	very fluid	5	not measured- specimens damaged during setting
10	2.	workable	12	7.6
	C 3.	fluid	6	not measured- specimens damaged during setting
15	C 4.	workable	6	1.9
	5.	workable	26	6.7
20	C 6.	workable	12	not measured- specimens damaged during setting
	C 7.	workable	7	6.4
25	8.	workable	47	6.5
	C 9.	fluid	15	not measured- specimens damaged during setting
30				

Comparative Examples 1, 4 and 7 evolved large amounts of ammonia gas upon setting, such that the strength specimens expanded out of their molds. Comparative Examples 3, 6 and 9 evolved ammonia gas during mixing, and for several minutes after the strength specimens were cast. This caused the material to flow out of the molds and then to shrink back, resulting in half-formed cubes of high porosity. The majority of the comparative samples were damaged by the gas evolution during their rapid set, such that the 24 hour compressive strength test could not be carried out on them.

These problems were not encountered with the binders of the Examples of the present invention, which set more slowly than the comparative materials though still quite rapidly, in less than about three quarters of an hour. The 24 hour compressive strengths of all the samples of the present invention were equivalent to or better than the best of the comparative samples which survived to testing.

A second series of tests were carried out as described above, to compare Examples of the present invention with prior art cementitious binders. In these tests, the amount of retarder was varied in order to attempt to match setting times for the different formulations. The binder formulations are set forth in Table 3 below.

**Table 3**

<u>Example</u>	<u>Binder Mixture Formulation</u>
C 10.	MgO(38%) - $\text{NH}_4\text{H}_2\text{PO}_4$ (43%) - $\text{H}_2\text{O}$ (15%) + boric acid (4%)
11.	MgO(29%) - $\text{KH}_2\text{PO}_4$ (55%) - $\text{H}_2\text{O}$ (16%)
C 12.	MgO(33%) - $\text{NH}_4\text{H}_2\text{PO}_4$ (31%) - $\text{K}_2\text{CO}_3$ (18%) - $\text{H}_2\text{O}$ (15%) + boric acid (2.5%)
13	MgO(28%) - $\text{KH}_2\text{PO}_4$ (55%) - $\text{H}_2\text{O}$ (16%) + boric acid (0.3%)
14.	MgO(28%) - $\text{KH}_2\text{PO}_4$ (55%) - $\text{H}_2\text{O}$ (16%) + boric acid (0.5%)
C 15.	MgO(37%) - $\text{NH}_4\text{H}_2\text{PO}_4$ (41%) - $\text{H}_2\text{O}$ (14%) + boric acid (8%)
C 16.	MgO(35%) - $\text{NH}_4\text{H}_2\text{PO}_4$ (39%) - $\text{H}_2\text{O}$ (14%) + boric acid (12.5%)
C 17.	MgO(34%) - $\text{NH}_4\text{H}_2\text{PO}_4$ (32%) - $\text{K}_2\text{CO}_3$ (18%) - $\text{H}_2\text{O}$ (15%) + boric acid (0.5%)
C 18.	MgO(34%) - $\text{NH}_4\text{H}_2\text{PO}_4$ (32%) - $\text{K}_2\text{CO}_3$ (18%) - $\text{H}_2\text{O}$ (15%) + boric acid (1%)

Again, the fluidities of the initial mixtures were evaluated, the setting times measured, and the compressive strengths of the molded articles tested after 24 hours. Results of the tests are set forth in Table 4, below.

5

**Table 4**

	<u>Example</u>	<u>Fluidity</u>	<u>Setting time</u> <u>minutes</u>	<u>24 hour compressive strength (MPa)</u>
10	C 10.	workable	17	10.9
	11.	workable	16	10.9
15	C 12.	fluid	58	30.3
	13.	workable	44	5.1
	14.	workable	44	5.2
20	C 15.	workable	29	19.6
	C 16.	workable	41	20.7
25	C 17.	workable	25	52.6
	C 18.	fluid	22	19.4

30

Comparative Examples 10, 15 and 16 evolved such large amounts of ammonia gas upon setting that the strength specimens expanded out of their molds.

35

In Comparative Example 12, the  $\text{NH}_4\text{H}_2\text{PO}_4$ , water, and  $\text{K}_2\text{CO}_3$  were allowed to interact for more than 30 minutes, such that gas evolution could occur before MgO and boric acid were added; gas was still evolved during the setting process, making preparation of specimens for strength testing difficult.

In Comparative Example 17, the  $\text{K}_2\text{CO}_3$  was ground to less than 500 microns, then  $\text{NH}_4\text{H}_2\text{PO}_4$ , water, and  $\text{K}_2\text{CO}_3$  were allowed to interact for about 30 minutes, such that gas evolution could occur before MgO and boric acid were added.

In Comparative Example 18, the  $K_2CO_3$  was dissolved in the mix water, then  $NH_4H_2PO_4$ , water, and  $K_2CO_3$  were allowed to interact for more than 30 minutes, such that gas evolution could occur before MgO and boric acid were added; gas was still evolved during the setting process, making preparation of specimens for strength testing difficult.

The problems associated with the evolution of ammonia gas were avoided in the Examples of the present invention, and acceptable setting times and compressive strengths were achieved.

The inventive cementitious binder is useful in mortar formulations which include substantially inert fillers, including but not limited to silica sand, Class F fly ash, talc, clay based sand, silica fume and mixtures thereof.

The inert filler may be added to the cementitious binder in an amount of about 60% to about 85%, preferably about 70% to about 80% based on the total weight of the resulting mortar to achieve the inventive mortar composition.

Cementitious mortars were prepared according to the method of the present invention and tested for initial set time, and both 24 hour and 7 day compressive strength. The formulations and test results are set forth below as Examples 19-24.

#### Example 19

Reactive powder (29% MgO, 60% $KH_2PO_4$ , 11% sodium tripolyphosphate):	553g
Fly ash:	392g
Silica sand:	2354g
Water	345mL

Initial set time = 18 minutes  
 24 hour compressive strength = 28.0 Mpa  
 7 day compressive strength = 29.8 Mpa

Example 20

5      Reactive powder  
       (21.8% MgO, 72%  $\text{KH}_2\text{PO}_4$ , 0.9% boric acid, 5.3% sodium tripolyphosphate):  
       Fly ash: 553g  
       Silica sand: 392g  
       Water: 2354g  
                   345mL

10      Initial set time = 38 minutes  
       24 hour compressive strength = 14.2 Mpa  
       7 day compressive strength = 25.5 Mpa

Example 21

15      Reactive powder  
       (45% MgO, 55%  $\text{KH}_2\text{PO}_4$ ):  
       Fly ash: 553g  
       Silica sand: 392g  
       Water: 2354g  
                   345mL

20      Initial set time = 24 minutes  
       24 hour compressive strength = 21.9 Mpa  
       7 day compressive strength = 22.0 Mpa

25

Example 22

30      Reactive powder  
       (45% MgO, 50%  $\text{KH}_2\text{PO}_4$ , 5% sodium tripolyphosphate):  
       Fly ash: 553g  
       Silica sand: 392g  
       Water: 2354g  
                   345mL

35      Initial set time = 14 minutes  
       24 hour compressive strength = 19.2 Mpa  
       7 day compressive strength = 25.1 Mpa

Example 23

40      Reactive powder  
       (33.2% MgO, 50%  $\text{KH}_2\text{PO}_4$ , 3.5% boric acid, 13.3% sodium tripolyphosphate):  
       Fly ash: 553g  
       Silica sand: 392g  
       Water: 2354g  
                   345mL

45

Initial set time = 37 minutes  
 24 hour compressive strength = 24.7 Mpa  
 7 day compressive strength = 35.1 Mpa

5      Example 24

	Reactive powder	
	(21.5% MgO, 50% KH <sub>2</sub> PO <sub>4</sub> , 3.5% boric acid, 25% sodium tripolyphosphate):	
		553g
10	Fly ash:	392g
	Silica sand:	2354g
	Water	345mL

15      Initial set time = 89 minutes  
          24 hour compressive strength = 4.1 Mpa  
          7 day compressive strength = 14.7 Mpa

20      The mortar of the present invention provides similar properties to mortars that are bound by struvite, namely quick-setting, high early strength, volume stability and excellent bond to most substrates. However, the potassium struvite mortar of the present invention eliminates the evolution of ammonia gas as a reaction by-product, gives a setting which is desirably slower but still rapid, and results in a lower heat evolution upon setting. These advantages permit the mortar to be used indoors, coated with a coating, and finished in place.

25

30      The cementitious composition mixtures of the present invention as described above may be formulated in dry form, to be activated by the introduction of water to the mixture. Accordingly, such a cementitious composition formulation would preferably comprise active ingredients consisting essentially of i) a compound capable of providing, in aqueous solution, magnesium ions, ii) a potassium phosphate compound, and iii) optionally an additional phosphate source; in which said mixture ammonia or an ammonium ion-containing compound is substantially absent. The active ingredients are preferably present in an amount of about 10 to about 35 parts of the magnesium ion-providing compound, about 30 to about 60 parts of the potassium phosphate compound, and up to about 20 parts of the additional phosphate source. The formulation may further comprise a substantially inert filler, and optionally a retarder.

35

Thus it is demonstrated that the objects of the present invention are met. The examples included above are for illustrative purposes only and the present invention is not to be limited to them. It is to be understood that other active ingredients, fillers, retarders, dispersants, accelerators and other admixtures and the like can be included in the cementitious materials of the present invention, without departing from the spirit of the invention herein disclosed and described. Thus, the scope of the invention shall include all modifications and variations that may fall within the scope of the attached claims and equivalent embodiments.

## CLAIMS

1. A method for making a cementitious material comprising:  
forming a mixture of water and active ingredients consisting essentially of i) a compound capable of providing, in aqueous solution, magnesium ions, ii) a potassium phosphate compound, and iii) optionally an additional phosphate source; wherein said mixture is formed substantially in the absence of ammonia or an ammonium ion-containing compound.
2. The method of claim 1 wherein said compound capable of providing magnesium ions is  $\text{MgO}$ , and said potassium phosphate compound is  $\text{KH}_2\text{PO}_4$ .
3. The method of claim 1 wherein forming the mixture includes adding a substantially inert filler.
4. The method of any one of claims 1-3 wherein forming the mixture includes adding a retarder.
5. The method of claim 1, wherein the active ingredients are present in an amount of about 10 to about 35 percent by weight of the magnesium ion-providing compound, and about 30 to about 60 percent by weight of the potassium phosphate compound, based on the weight of the water and the active ingredients.
6. A cementitious binder prepared by the method of claim 1 or 2.
7. A mortar prepared by the method of claim 3.
8. A method for repairing a flaw in a cementitious structure, comprising applying to the flaw a cementitious binder prepared by  
forming a mixture of water and active ingredients consisting essentially of i) a compound capable of providing, in aqueous solution, magnesium ions, ii) a potassium



phosphate compound, and iii) optionally an additional phosphate source;

wherein forming said mixture is substantially in the absence of ammonia or an ammonium ion-containing compound.

9. A method for repairing a flaw in a cementitious structure, comprising applying to the flaw a mortar prepared by

forming a mixture of an inert filler, water, and active ingredients consisting essentially of i) a compound capable of providing, in aqueous solution, magnesium ions, ii) a potassium phosphate compound, and iii) optionally an additional phosphate source;

wherein said mixture is formed substantially in the absence of ammonia or an ammonium ion-containing compound.

10. A cementitious composition formulation comprising active ingredients consisting essentially of i) a compound capable of providing, in aqueous solution, magnesium ions, ii) a potassium phosphate compound, and iii) optionally an additional phosphate source; in which said mixture ammonia or an ammonium ion-containing compound is substantially absent,

wherein the active ingredients are present in an amount of about 10 to about 35 parts of the magnesium ion-providing compound, about 30 to about 60 parts of the potassium phosphate compound, and up to about 20 parts of the additional phosphate source.



# The Patent Office

17

Application No: GB 9820167.6  
Claims searched: 1 to 10

Examiner: Miss M M Kelman  
Date of search: 2 December 1998

## Patents Act 1977 Search Report under Section 17

### Databases searched:

UK Patent Office collections, including GB, EP, WO & US patent specifications, in:

UK CI (Ed.P): C1H HBX

Int CI (Ed.6): C04B 9/00, 9/04, 12/00, 12/02, 28/00, 28/10, 28/34

Other: ONLINE: EDOC, JAPIO, WPI

### Documents considered to be relevant:

Category	Identity of document and relevant passage	Relevant to claims
X	WO 97/34848 A1 THE UNIVERSITY OF CHICAGO see page 11, line 28, to page 12, line 6, and Example 5	1,2,5,6,7,10
X	WO 97/21639 A1 RHONE-POULENC see the claims, page 3, lines 6 to 7, and line 17 to page 4, line 7, page 6, line 5, to page 7, line 10, and page 8, lines 30 to 38	1,2,3,4,6,7,8,9
X	US 5595597 A RHONE-POULENC see the claims, column 1, lines 15 to 25, column 2, line 30, to column 3, line 59, and column 4, lines 20 to 33	1,2,3,4,6,7,8,9

X Document indicating lack of novelty or inventive step  
Y Document indicating lack of inventive step if combined with one or more other documents of same category.  
& Member of the same patent family

A Document indicating technological background and/or state of the art.  
P Document published on or after the declared priority date but before the filing date of this invention.  
E Patent document published on or after, but with priority date earlier than, the filing date of this application.